

Nickel element doping impacts on structure features and Faraday effects of magneto-optical transparent holmium oxide ceramics

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Abstract

The nickel element doped holmium oxide ($\text{Ho}_2\text{O}_3:\text{Ni}$) transparent magneto-optical ceramics were fabricated by vacuum sintering and the dopant impacts on structure features and Faraday effects were investigated. The starting oxide powders were synthesized by pyrolyzing the resulting layered holmium-based hydroxide nanosheets prepared from a chemical precipitation route using the sodium hydroxide as precipitant at the freezing temperature. Upon high-temperature sintering, the $\text{Ni}_i^{\cdot\cdot}$ defect is introduced by Ni^{2+} substitution for Ho^{3+} to form the interstitial solid solution. The 1 at.% Ni^{2+} doped Ho_2O_3 ceramic sample exhibits an in-line transmittance of $\sim 70.04\%$ at 1550 nm with a relative density of $\sim 99.88\%$, while more Ni^{2+} incorporation (e.g., 2–5 at.%) even leads to a completely opaque state. The magneto-optical transparent $\text{Ho}_2\text{O}_3:1\%\text{Ni}$ ceramic developed in this work has Verdet constants of ~ -195 , -65 , and -29 rad/(T·m) at 635, 1064, and 1550 nm, respectively, which are ~ 1.8 -fold higher than the commercial terbium gallium garnet crystal or ~ 1.4 -fold higher than the pure Ho_2O_3 ceramic. This material also possesses relatively large figure of merit of ~ 14.6 °/T at 1064 nm and relatively high thermal conductivity of ~ 7.5 W/(m·K) at room temperature.

Keywords: magneto optics, transparent ceramics, rare earths, defects, Faraday effect

1. INTRODUCTION

The Faraday effect refers to a magneto-optical phenomenon caused by alterations in the magnetic field, which induces the change of the transmission characteristics of the incident light beam inside a transparent material.¹⁻³ Based on it, a magnetically-activated material can be employed as a magneto-optical element (MOE) in systems such as Faraday isolators, Faraday modulators, and magneto-optical switches.⁴⁻⁶ It can also indirectly monitor the strength of the magnetic field.⁶ Common magneto-optical materials can be classified as films, glass, crystals, and ceramics. At present, the most widely used magneto-optical medium is the terbium gallium garnet (TGG) single crystal with a Verdet constant of $-134 \text{ rad}/(\text{T}\cdot\text{m})$ at 633 nm .⁷ According to the Faraday effect, the Faraday rotation angle is positively related to the size and magnetic induction. Therefore, a large-size TGG crystal is required to enhance the Faraday effect so as to obtain a large Faraday rotation angle, which may generate more light loss and accumulated heat to possibly induce depolarization effect and thermal lens effect.^{8,9} In order to overcome this problem, more and more researchers recently devoted to developing new-type magneto-optical materials with a higher Verdet constant. The most studied magneto-optical ceramics primarily include lanthanide sesquioxides (Ln_2O_3), garnets, pyrochlores, and fluorites,¹⁰⁻¹⁶ among which the sesquioxides generally possess the largest Verdet constant due to the highest Ln^{3+} concentration. In addition, a regular strategy for enhanced Verdet constant in ceramic system is to being doped with another Ln^{3+} cation such as Pr^{3+} , Ce^{3+} , Nd^{3+} , Ho^{3+} , Dy^{3+} , or Tm^{3+} via super-exchange interaction.^{3,17-20} However, these Ln^{3+} cations contain abundant energy levels, frequently causing more optical self-absorption to limit the practical application.

According to the Hund rule and Pauli exclusion principal, electrons are always arranged in pairs with opposite spins. The magnetic moments produced by two

electrons with opposite spin directions would cancel each other out. Thus, unpaired electrons are desired in the outermost layer of the nucleus for achieving magnetic moments. The main elements that meet this condition are Fe, Co, and Ni, among which the Fe element has variable valence while the abundance of the Co element is much less than that of the Ni element. Furthermore, the Ni element has good resistance corrosion and chemical stability. NiO is a typical natural antiferromagnet,^{21,22} however, Ho₂O₃ is paramagnetic. Their combined influence on Faraday magneto-optical effect has scarce study so far. In this work, we fabricated transparent Ni²⁺ doped Ho₂O₃ ceramics and studied dopant impacts on structure features and Verdet constants.

In order to obtain a sinterable starting powder, we herein prepared exchangeable layered rare-earth hydroxide (LRH) as the precipitation precursor using a freezing temperature technique, since the obtained LRH nanosheet down to ~7 nm has proved its superiority for the production of high-optical-quality oxide ceramic.²³⁻²⁶ During the liquid-phase synthesis, the ammonium hydroxide solution is frequently employed as the precipitant. However, such a precipitant is unable to precipitate nickel cation rather than preference for the formation of soluble coordination compound, especially under a high pH level. Considering the sodium hydroxide solution also could provide adequate OH⁻ anions for LRH synthesis and the soluble Na⁺ byproduct could be removed by subsequent washing. Hence, we alternatively utilized the sodium hydroxide solution as the precipitant in the present work, finally leading to transparent Ho₂O₃:Ni ceramics through vacuum sintering. More importantly, unlike the Ln³⁺ cation, the Ni²⁺ dopant never generates additional optical self-absorption in the near-infrared region, which avoids imposing restriction on the application of magneto-optical Ho₂O₃ ceramics. The research findings of this work could provide a reference for the development of high-performance magneto-optical materials.

2. EXPERIMENTAL PROCEDURES

The Ho_2O_3 raw material (99.95% purity, Shanghai Diyang Chemical Co., Ltd., Shanghai, China) was dissolved by excessive amount of nitric acid upon heating. The superfluous nitric acid was fully removed by evaporating the salt solution until dryness. The dried nitrate was dissolved into distilled water to prepare a 0.075 mol/L $\text{Ho}(\text{NO}_3)_3$ mother liquor. The stoichiometric $\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$ (99.9% purity, Aladdin Biochemical Technology Co., Ltd., Shanghai, China) was dissolved into the mother liquor according to 1–5 at.% cationic concentration. Meanwhile, a 1.0 mol/L NaOH solution was prepared by dissolving its commercial solid powder (98% purity, Shanghai Macklin Biochemical Co., Ltd.) into distilled water as the precipitant. The alkaline precipitant was dropwise dropped into the mother liquor at a rate of 2–3 mL/min under magnetic stirring at a freezing temperature of $\sim 4^\circ\text{C}$ till the pH value reached 8.5. The resulting light green suspension, after aging for 1.5 h, was filtered and repeatedly washed with distilled water to remove byproducts. The precipitation was added into $(\text{NH}_4)_2\text{SO}_4$ solution and the molar ratio of SO_4^{2-} to Ho^{3+} was selected to be 0.03. After ion exchange for 1.5 h, the product was repetitively rinsed with distilled water and absolute ethanol. After drying at 90°C for 12 h, the precipitation precursor was calcined in a tube furnace under flowing oxygen (~ 150 mL/min) at 1050°C for 4 h to yield the oxide particle. The powder was filled into a stainless-steel die for precompression, and then cold isostatically pressed under 300 MPa. The green body was sintered in a tungsten-wire furnace at 1850°C for 6 h under 10^{-4} – 10^{-5} Pa vacuum. The sintered body was finally polished on double faces to improve the surface smooth finish.

Phase identification was performed on X-ray diffractometer (XRD; Model D8 Advance Davinci, Bruker, Karlsruhe, Germany) using nickel filtered Cu $K\alpha$ radiation. The particle size distribution was determined by laser diffraction particle size analyzer

(Model ZS90, Malvern Instruments, Malvern, UK). The microscopic morphologies of the particle and ceramic samples were both observed on a field emission scanning electron microscope (FE-SEM; Model Nova NanoSEM 450, FEI Company, Hillsboro, USA) at 15–30 kV, which is equipped with an Apollo silicon drift detector series for energy-dispersive X-ray spectrometry (EDS). The in-line transmittances of sintered bodies were recorded on a UV/VIS/NIR spectrophotometer (Model Lambda 950, PerkinElmer, Shelton, USA) from 200 to 1700 nm. The Faraday rotation angles were measured using the extinction method on a self-assembled device equipped with three independent laser light sources (635, 1064, and 1550 nm) under adjustable magnetic field from 0 to 0.8 T. Thermal diffusivity and specific heat value were simultaneously obtained from a laser flash diffusivity apparatus (Model LFA467, Netzsch, Selb, Germany) using xenon lamp as the laser heating source from room temperature to 400 °C.

3. RESULTS AND DISCUSSION

Figure 1 shows XRD patterns of the precipitation precursor, calcination product, and sintered bodies with 1–5 at.% Ni²⁺ doping. The precursor exhibits typical LRH structure with characteristic sharp (220) diffraction and a series of (00*l*) (*l* = 2, 4···) diffractions.²⁷ The lattice constant (*c*) of LRH is calculated to be ~1.822 nm using Bragg's equation from the (002) diffraction. The interlayer distance for LRH is half of the *c* constant,²⁸ and thus its value is ~0.911 nm. After calcining at 1050 °C for 4 h, the diffraction peaks of the product match well with those of the Ho₂O₃ standard card (JCPDS No. 83-0932) and the sharp diffraction peaks indicate the high crystallinity. Owing to the close similarities between (001) direction for LRH and (111) direction for oxide, the phase transformation substantially belongs to quasi-topotactic evolution.²⁹ The crystallite size (*D*_{XRD}) can be calculated from the full width at half maximum

(FWHM) of the (222) diffraction band using Scherrer's equation: $D_{XRD} = K\lambda/(\beta\cos\theta)$, where K is the shape factor ($K = 0.89$), λ is the wavelength of X-ray ($\lambda = 0.15406$ nm), β is the FWHM, and θ is the Bragg angle.^{30,31} The determined crystallite size of the Ho₂O₃:1%Ni oxide powder is ~39 nm.

After high-temperature sintering, the three Ho₂O₃ bulk specimens with 1–5 at.% Ni²⁺ doping still retain the cubic crystalline phase and no trace of impurities can be observed. The main (222) diffraction peaks gradually shift towards the low angle side with more Ni²⁺ addition. The lattice constants calculated by Bragg's law are ~10.572, 10.586, and 10.592 Å for 1, 2, and 5 at.% Ni²⁺ doped Ho₂O₃ samples, respectively. These phenomena indicate that homogeneous solid solutions have formed. Ho₂O₃ belongs to cubic C-type sesquioxide, which has 80 atoms per unit cell containing 32 Ho atoms and 48 O atoms. Its X-ray density (d_{th}) could be determined from Eq. (1):

$$d_{th} = \frac{32 \times [(1-m)M_{Ho} + mM_{Ni} + 1.5M_O]}{a^3 N_A} \quad (1)$$

where M_i stands for atomic weight of element i ($i = \text{Ho, Ni, and O}$), a refers to the lattice constant, N_A represents the Avogadro constant, and m denotes the percentage content of Ni atom. Their corresponding theoretical densities are ~8.46, 8.42, and 8.40 g/cm³. The slightly decreasing theoretical density is because Ni element is lighter relative to Ho element while the increasing lattice volume with more Ni²⁺ doping is another one reason.

The Ni²⁺ doping mechanism is discussed as follows: Since the oxidation states of Ni and Ho elements differ (Ni being 2+ and Ho being 3+), the introduction of low-valent Ni²⁺ may induce one of two possible point defect forms, as described in Eqs. (2) and (3):





In the first case, every two divalent Ni acceptors create one $V_O^{\cdot\cdot}$ (Eq. 2), while in the second case, the interstitial $Ni_i^{\cdot\cdot}$ cation is formed (Eq. 3). The Ni^{2+} ionic radius [0.690 Å for coordination number (CN) = 6] is smaller than that of Ho^{3+} (0.901 Å for CN = 6). The formation of oxygen vacancies, as described in Eq. (2), results in the contraction of the unit cell. However, this was inconsistent with the XRD results. Sometimes, an exaggerated mismatch in ionic radii could lead to the formation of an interstitial solid solution according to Eq. (3). Thus, the smaller Ni^{2+} penetrates the Ho_2O_3 lattice, expanding the cell dimensions. There are two inequivalent Ho^{3+} sites in Ho_2O_3 unit cell. Either Ho^{3+} is bonded to six equivalent O^{2-} atoms to form a mixture of distorted edge and corner-sharing HoO_6 octahedra or to form a mixture of edge and corner-sharing HoO_6 octahedra. Overall, the Ni^{2+} dopant cannot form substitutional solid solution, and thus Ni^{2+} site is impossible at the octahedral center. The formation of the interstitial $Ni_i^{\cdot\cdot}$ defect in the Ho_2O_3 lattice was confirmed by aforementioned analysis.

Figure 2 exhibits the morphologies of the precipitation precursor synthesized at the freezing temperature and the Ho_2O_3 :1%Ni oxide powder calcined at 1050 °C for 4 h. The LRH precursor presents typical two-dimensional nanosheet shape and self-assembles into collective three-dimensional petal-like pattern. After thermal decomposition, the nanosheets collapse into rounded powders with ultrafine size. The average particle size measured by laser diffraction particle size analyzer is ~263 nm [inset in Fig. 2(b)], while the statistic FE-SEM size is ~83 nm. The former is normally larger than the latter for the non-monodispersed powder due to the self-limitativeness of laser diffraction measurement method.³² The result of the laser diffraction particle size analyzer could partly reflect the agglomerated extent. Our determined particle size test result is relatively small in fact while the particle size curve in differential volume

distribution exhibits a unimodal particle size distribution without observed hard agglomerates, implying the good particle sinterability for dense ceramic preparation.

Figure 3(a) shows the microstructure of the Ho₂O₃:5%Ni bulk sample, especially focusing on the triple junction region. The grain boundary can be clearly observed after thermal etching treatment and the element mapping results reveal that the elemental distributions of holmium, oxygen, and nickel are relatively uniform without detected secondary phases [Figs. 3(b)-(e)]. That is, the Ni²⁺ dopant has well dissolved into the Ho₂O₃ matrix. The elemental distribution analysis is fully consistent with the XRD result [Fig. (1)]. The qualitative EDS analysis also detects Ho, Ni, and O elements, whereas the additional Pt element ascribes to the surface sputtering for electrical conduction [Fig. 3(f)]. By excluding the Pt-coating effect, the semiquantitative EDS analysis reveals that the elemental contents of Ho, Ni, and O respectively are 74.09, 5.19, and 20.72 at.% in close proximity to our practical stoichiometric proportion, which further verifies that our wet chemical route is feasible for Ni²⁺ and Ho³⁺ precipitation completeness.

Figure 4(a) exhibits appearances of the Ho₂O₃:xNi²⁺ ($x = 1-5$ at.%) ceramic bodies fabricated by vacuum sintering. The letters covered under the Ho₂O₃:1%Ni bulk can be well read-through. This transparent body presents brown hue similar to the pure Ho₂O₃ ceramic,¹³ since the intra-4f¹⁰ transitions for Ho³⁺ greatly absorb the purple, blue, and green lights in the visible region [Fig. 4(b)]. Additionally, the absorption bands arising from 3d-3d transitions of Ni²⁺ in the visible region almost overlap with the abundant intra-4f¹⁰ transitions of Ho³⁺ to be difficult to separately distinguish. Based on Tanabe-Sugano diagram, only ³A₂→³T₂ transition of Ni²⁺ may generate self-absorption effect in the infrared region, but it has been swallowed in the Ho³⁺ ⁵I₈→⁵I₆ transition induced absorption band. That is to say, the Ni²⁺ introduction never generates additional

individual optical absorption in the infrared region, indicating its merit for application in this spectrum scope. The overall transparency has not been reduced by 1 at.% Ni²⁺ doping in comparison with our previously reported pure Ho₂O₃ ceramic.¹³ Nevertheless, the color of the samples gradually deepens as the increase in Ni²⁺ contents, even leading to completely opaque black close to NiO own color. The theoretical transmittance of the defect-free Ho₂O₃ single crystal has been reported to be ~81.6% at 1550 nm.²⁵ In this work, the Ho₂O₃:1%Ni sample has a transparency of ~70.0% at the same wavelength, which is ~86% of the theoretical value. Unlike the case in the visible region, the near-infrared range has much less absorption from electronic transitions, and thus this material exhibits good potential to apply to the high-power near-infrared laser system such as Yb³⁺ or Nd³⁺ doped Y₃Al₅O₁₂ laser device. The density of the Ho₂O₃:1%Ni ceramic has been measured by the Archimedes method and the result shows that this specimen has an experimental density of ~8.45 g/cm³. Considering its theoretical density is ~8.46 g/cm³ (Eq. 1), the relative density is as high as ~99.88% in accordance with the good optical quality.

A self-assembled magneto-optical test equipment is used to measure the Faraday effect of the transparent Ho₂O₃:1%Ni ceramic sample via the extinction method. In this system, three lasers with wavelengths of 635, 1064, and 1550 nm are used as light sources, and the magnetic field intensity is adjusted from 0 to 0.8 T through current-controlled magnet coil. The relationship between Verdet constant (V) and Faraday rotation angle (θ) can be expressed as: $V = \theta / BL$, where B is the applied magnetic field strength and L is the thickness of the transparent material.³³ Thus, a plot of B against θ/L has a linear correlation, where the slope (s) of the fitting is equal to the V parameter. The measured Faraday rotation angle exhibits a good linear relationship with the applied magnetic field [Fig. 5(a)]. The Verdet constants are determined to be ~-195, -

65, and -29 rad/(T·m) at 635, 1064, and 1550 nm, respectively. Table 1 compares the Verdet constants among our developed Ho₂O₃:Ni ceramics and several typical magneto-optical materials.^{13,25,34-39} At 1064 nm, the Verdet constant of our sample is ~1.8-fold higher than that of the commercial TGG crystal or ~1.4-fold higher than that of the pure Ho₂O₃ ceramic. Our results demonstrate that the nickel element is a quite effective dopant for enhanced Verdet constant via the super-exchange interaction. When the temperature remains constant, the V value and the test wavelength (λ) satisfy the following formula:

$$V = \frac{E}{(\lambda^2 - \lambda_0^2)} \quad (4)$$

In Eq. (4), λ_0 is relative with the effective migration wavelength of electrons and E is a constant.³⁴ A plot of V^{-1} versus λ^2 yields a good linear relation with a high degree of fitting ($R^2 = 99.94\%$) as shown in Fig. 5(b). The E and λ_0 parameters are determined to be $\sim 6.8 \times 10^7$ and 232 nm from the slope (E^{-1}) and the intercept ($-\lambda_0^2/E$), respectively.

Figure of merit (FoM), defined by the ratio of the Verdet constant to the light absorption coefficient (α), was used to roughly access the quality of a magneto-optical material as shown from Eqs. (5) and (6):

$$\text{FoM} = \frac{|V|}{\alpha} \quad (5)$$

$$\alpha = \frac{1}{L} \lg\left(\frac{1}{T}\right) \quad (6)$$

where T is the transmittance and L is the material thickness.^{35,40,41} The calculated FoM values are ~9.8, 14.6, and 7.3 °/T at 635, 1064, and 1550 nm, respectively. The lower FoM values at 635 and 1550 nm ascribe to the strong absorption from f - f transitions of Ho³⁺ and longer wavelength based on Eq. (4), respectively. At 1064 nm, the FoM value of our sample is much higher than those of the previously reported undoped Ho₂O₃

ceramic and the commercial TGG single crystal (Table 1).

The measured thermal diffusivity of the Ho₂O₃:1%Ni sample is shown in Fig. 6. The thermal diffusion (α) values are $\sim 2.94, 2.46, 2.07, 1.80,$ and $1.60 \text{ mm}^2/\text{s}$ at 25, 100, 200, 300, and 400 °C, respectively. The corresponding specific heats (C_p) are $\sim 0.302, 0.319, 0.347, 0.347,$ and $0.366 \text{ J}/(\text{g}\cdot\text{K})$. The thermal conductivity (k) can be deduced from the equation: $k = \alpha \cdot C_p \cdot \rho$, where ρ is the material density ($\sim 8.45 \text{ g}/\text{cm}^3$ for Ho₂O₃:1%Ni). As a result, the k values are $\sim 7.5, 6.6, 6.1, 5.3,$ and $5.0 \text{ W}/(\text{m}\cdot\text{K})$ at 25, 100, 200, 300, and 400 °C, respectively. As the ambient temperature rises, the thermal diffusion together with the thermal conductivity goes a declining trend. This is because the accelerated phonon vibration frequency and the increasing collision probability caused by elevated temperature shorten phonon mean free path. At room temperature, the thermal conductivity of our sample is much higher than those of previously reported Ho₂O₃:5%La ceramic [$\sim 4.8 \text{ W}/(\text{m}\cdot\text{K})$] and the commercial TGG crystal [$\sim 4.9 \text{ W}/(\text{m}\cdot\text{K})$],^{42,43} further suggesting the good application prospect as Faraday devices.

4. CONCLUSION

Layered holmium-based hydroxide nanosheets were prepared by wet-chemical method upon thermolysis into rounded Ho₂O₃:Ni oxide powders with an average particle size of $\sim 83 \text{ nm}$. The optimum Ni²⁺ concentration is 1 at.%, at which the sintered body exhibits an in-line transmittance of $\sim 70.04\%$ at 1550 nm ($\sim 86\%$ of the theoretical transmittance) and a thermal conductivity of $\sim 7.5 \text{ W}/(\text{m}\cdot\text{K})$ at room temperature. The nickel dopant in Ho₂O₃ has been proved to be effective to enhance the magneto-optical effect, by which the Ho₂O₃:1%Ni ceramic has ~ 1.4 -fold higher Verdet constant than the undoped Ho₂O₃ ceramic or ~ 1.8 -fold higher Verdet constant than the commercial TGG crystal. Our developed ceramic materials exhibit good potential to apply to the high-power laser system, especially for use in the near-infrared region.

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CONFLICT OF INTEREST STATEMENT

The authors declare they have no conflicts of interest.

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Figure captions

Fig. 1. XRD patterns of the precipitation precursor intended for $\text{Ho}_2\text{O}_3:1\%\text{Ni}$, the calcination product at $1050\text{ }^\circ\text{C}$, and the sintered bodies with 1–5 at.% Ni^{2+} doping for comparison. The inset in the right-hand panel shows the enlarged view of the main (222) diffraction peaks.

Fig. 2. FE-SEM micrographs of the LRH precursor intended for $\text{Ho}_2\text{O}_3:1\%\text{Ni}$ (a) and its calcined powder at $1050\text{ }^\circ\text{C}$ for 4 h (b).

Fig. 3. FE-SEM micrograph of the $\text{Ho}_2\text{O}_3:5\%\text{Ni}$ ceramic sample (a), elemental mappings for Ho (b), O (c), Ni (d) and their comprehensive distributions (e), and EDS elemental analysis (f).

Fig. 4. Appearances (a) and in-line transmittances (b) of the $\text{Ho}_2\text{O}_3:x\text{Ni}^{2+}$ ($x = 1\text{--}5$ at.%) ceramics fabricated by vacuum sintering.

Fig. 5. A plot of θ/L versus B for the transparent magneto-optical Ho_2O_3 ceramic doped with 1 at.% Ni^{2+} (a) and a plot of $1/V$ versus λ^2 for the $\text{Ho}_2\text{O}_3:1\%\text{Ni}$ bulk sample (b).

Fig. 6. Thermal diffusivity and thermal conductivity of the transparent $\text{Ho}_2\text{O}_3:1\%\text{Ni}$ ceramic.